



Equations Practice

Write net ionic equations for each of the following. Each of the reactions occurs in aqueous solution unless otherwise indicated. Represent substances in solution as ions if the substance is extensively ionized. Omit formulas for any ions or molecules that are unchanged by the reaction. In all cases a reaction occurs. Balance all equations.

1. Solid magnesium oxide is added to distilled water.



2. Solid potassium hydride is added to distilled water.



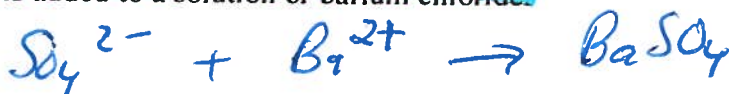
3. Solutions of silver nitrate and sodium phosphate are mixed.



4. Solid potassium hydrogen carbonate is added to aqueous hydrochloric acid.



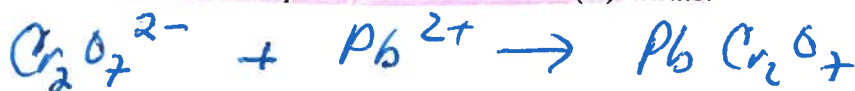
5. Dilute sulfuric acid is added to a solution of barium chloride.



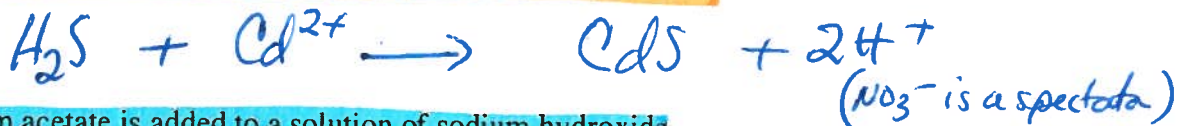
6. Solid lithium chlorate is heated.



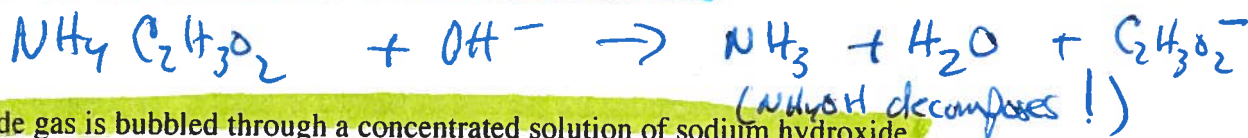
7. Solid potassium dichromate is added to an aqueous solution of lead(II) nitrate.



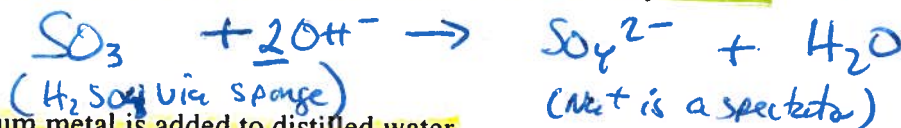
8. Hydrogen sulfide gas is bubbled through a solution of cadmium nitrate.



9. Solid ammonium acetate is added to a solution of sodium hydroxide.



10. Sulfur trioxide gas is bubbled through a concentrated solution of sodium hydroxide.



11. A small piece of potassium metal is added to distilled water.

